↳ chemical construction of water :

weight as [redacted]

$H_2O$

% 11,11 ← H O → % 88,89

| | H | O |
|---|---|---|
| as ratio of atoms | 2 | 1 |
| as mass | 1 | 8 |

↳ the bond between H and O is covalent bond
↳ the angel between H and O is obtuse angel
= 104,5°

Electronegativity :
↳ when atom attract electrons
↳ So when atom attract more electrons it called The most negative and if it attract small number of electrons it calle the least [redacted] negative

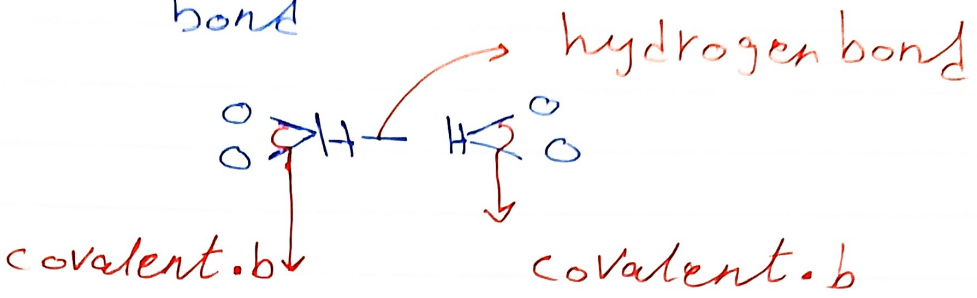
↳ when atom get electron It called Negative Ion ($Cl^-$)
↳ when atom lost electron it called positive Ion ($Na^+$).

↳ polar bond : bond between two negative Ions
↳ the atom which has more electrons ($\delta^-$), but the atom which has less electrons ($\delta^+$)

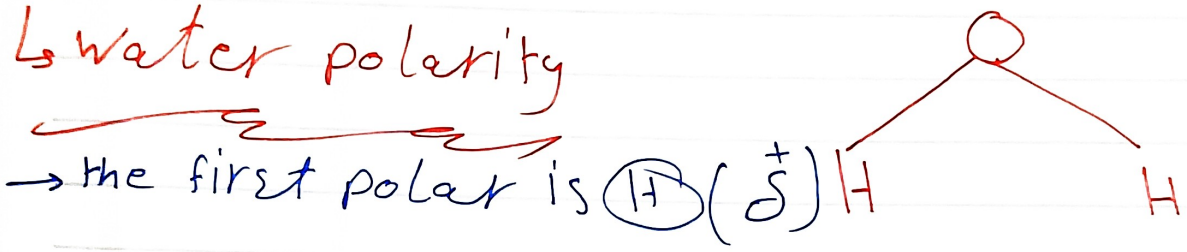
↳ Polar molecule : Molecule contain 2 charges one of them + and the another −

$^+H_2O^-$

↳ hydrogen bond: bond between 2 hydrogen atoms in two different polar bond




↳ water polarity



→ the first polar is H ($\delta^+$)

→ the second polar is O ($\delta^-$)

→ there is 2 polar bonds between O and H becaus the oxgen is more electronegativity Than hydrogen

→ Due to the polarity of water, it can form bond betwee water molecules and anothe compound, So it can dissolve any salt to ions

$$NaCl + H_2O \longrightarrow Na^+ Cl^-$$

water broke the bonds between Nacl and covered This ions to preven connect again.

→ The boilling points of $H_2O$ more than $H_2O$ due to the hydrogen bond between H and O (4) more The s and o