

## Unit 1 : lesson 2 : Periodic table for classification of elements

Chemical elements differ in physical properties and chemical properties... Therefore, there have been multiple attempts to classify the elements, and their goal is to facilitate study and find the relationship between the elements and their physical and chemical properties.

### The most important attempts to classify elements:-

1) The periodic table of the scientist Mendeleev

2) The periodic table of the scientist Moseley

3) The modern periodic table

### 1) The periodic table of the scientist Mendeleev...

- The scientist Mendeleev's table is considered the first true periodic table for classifying elements...
- The elements are arranged in ascending order according to their atomic mass, without a regular progression when moving from left to right in horizontal rows.
- He discovered that the properties of the elements are repeated periodically at the beginning of each new row.. This means that the elements that exist in the form of vertical columns are similar in properties

Periodic Table of Elements  
based on Mendeleev's Periodic Law

0	I	II	III	IV	V	VI	VII	VIII			
He 4.00	H 1.01	Li 6.94	Be 9.01	B 10.8	C 12.0	N 14.0	O 16.0	F 19.0			
Ne 20.2	Na 23.0	Mg 24.3	Al 27.0	Si 28.1	P 31.0	S 32.1	Cl 35.5				
Ar 40.0	K 39.1	Ca 40.1	Sc 45.0	Ti 47.9	V 50.9	Cr 52.0	Mn 54.9	Fe 55.9	Co 58.9	Ni 58.7	
	•Cu 63.5	Zn 65.4	Ga 69.7	Ge 72.6	As 74.9	Se 79.0	Br 79.9				
Kr 83.8	Rb 85.5	Sr 87.6	Y 88.9	Zr 91.2	Nb 92.9	Mo 95.9	Tc (99)	Ru 101	Rh 103	Pd 106	
	•Ag 108	Cd 112	In 115	•Sn 119	Sb 122	Te 128	I 127				
Xe 131	Ce 133	Ba 137	La 139	Hf 179	Ta 181	W 184	Re 186	Os 194	Ir 192	Pt 195	
	•Au 197	•Hg 201	Tl 204	•Pb 207	Bi 209	Po (210)	At (210)				
Rn (222)	Fr (223)	Ra (226)	•Ac (227)	•Th 232	•Pa (231)	•U 238					

Legend:  
• Lanthanide series  
• Actinide series  
• Known to Ancients  
• Known to Mendeleev  
• Dobereiner's triads

## 2) The periodic table of the scientist Moseley...

- After the discovery of protons by the scientist Rutherford, the scientist Moseley discovered that the periodicity of the properties of elements is related to their atomic numbers and not to the atomic mass.
- Moseley modified Mendeleev's table in which the elements were arranged in ascending order according to their atomic number.
- Each element exceeds the element that precedes it in the same period by one integer
- Moseley added to the scientist Mendeleev's table the group of inert gases and other new elements that were discovered after the publication of Mendeleev's table.

1		2					3		4		5		6		7		0
H																	He
Li	Be											B	C	N	O	F	Ne
Na	Mg											Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra	Ac															

## 3) The modern periodic table:-

- The elements are arranged in ascending order according to their atomic numbers and the method of filling the sub-energy levels with electrons.

**Periodic Table of the Elements**

1 1IA 11A																	18 VIIIA 8A
1 H Hydrogen 1.0079	2 IIA 2A											13 IIIA 3A	14 IVA 4A	15 VA 5A	16 VIA 6A	17 VIIA 7A	2 He Helium 4.00260
3 Li Lithium 6.941	4 Be Beryllium 9.01218											5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.00674	8 O Oxygen 15.9994	9 F Fluorine 18.998403	10 Ne Neon 20.1797
11 Na Sodium 22.989768	12 Mg Magnesium 24.305	3 IIIB 3B	4 IVB 4B	5 VB 5B	6 VIB 6B	7 VIIB 7B	8 VIII 8	9 VIII 8	10 VIII 8	11 IB 1B	12 IIB 2B	13 Al Aluminum 26.981539	14 Si Silicon 28.0855	15 P Phosphorus 30.973762	16 S Sulfur 32.065	17 Cl Chlorine 35.4527	18 Ar Argon 39.948
19 K Potassium 39.0983	20 Ca Calcium 40.078	21 Sc Scandium 44.95591	22 Ti Titanium 47.88	23 V Vanadium 50.9415	24 Cr Chromium 51.9961	25 Mn Manganese 54.938	26 Fe Iron 55.847	27 Co Cobalt 58.9332	28 Ni Nickel 58.6934	29 Cu Copper 63.546	30 Zn Zinc 65.39	31 Ga Gallium 69.723	32 Ge Germanium 72.64	33 As Arsenic 74.92159	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80
37 Rb Rubidium 85.4678	38 Sr Strontium 87.62	39 Y Yttrium 88.90585	40 Zr Zirconium 91.224	41 Nb Niobium 92.90638	42 Mo Molybdenum 95.94	43 Tc Technetium 98.9072	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.9055	46 Pd Palladium 106.42	47 Ag Silver 107.8682	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.760	52 Te Tellurium 127.6	53 I Iodine 126.90447	54 Xe Xenon 131.29
55 Cs Cesium 132.90545	56 Ba Barium 137.327	57-71 Lanthanide Series	72 Hf Hafnium 178.49	73 Ta Tantalum 180.9479	74 W Tungsten 183.85	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.96655	80 Hg Mercury 200.59	81 Tl Thallium 204.3833	82 Pb Lead 207.2	83 Bi Bismuth 208.98037	84 Po Polonium [209]	85 At Astatine [209]	86 Rn Radon [222]
87 Fr Francium 223.0197	88 Ra Radium 226.0254	89-103 Actinide Series	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [269]	109 Mt Meitnerium [268]	110 Ds Darmstadtium [271]	111 Rg Roentgenium [272]	112 Cn Copernicium [285]	113 Uut Ununtrium [284]	114 Fl Flerovium [289]	115 Uup Ununpentium [288]	116 Lv Livermorium [293]	117 Uus Ununseptium [294]	118 Uuo Ununoctium [294]
		57 La Lanthanum 138.9055	58 Ce Cerium 140.116	59 Pr Praseodymium 140.90765	60 Nd Neodymium 144.24	61 Pm Promethium 144.9127	62 Sm Samarium 150.36	63 Eu Europium 151.9654	64 Gd Gadolinium 157.25	65 Tb Terbium 158.92534	66 Dy Dysprosium 162.50	67 Ho Holmium 164.93032	68 Er Erbium 167.26	69 Tm Thulium 168.93421	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967	
		89 Ac Actinium 227.0278	90 Th Thorium 232.0381	91 Pa Protactinium 231.03688	92 U Uranium 238.02891	93 Np Neptunium 237.0482	94 Pu Plutonium 244.0642	95 Am Americium 243.0614	96 Cm Curium 247.0703	97 Bk Berkelium 247.0703	98 Cf Californium 251.0798	99 Es Einsteinium [254]	100 Fm Fermium [257]	101 Md Mendelevium [258]	102 No Nobelium [259]	103 Lr Lawrencium [262]	
		Alkali Metals	Alkaline Earths	Transition Metals	Basic Metals	Semi-Metals	Nonmetals	Halogens	Noble Gases	Lanthanides	Actinides						

## Description of the modern periodic table:-

- The modern periodic table consists of 7 horizontal period and 18 vertical groups

( 7 cycles because the number of energy levels is 7 levels... In each cycle we begin to fill a new energy level with electrons )

- The number of elements in the modern periodic table is 118

Elements are found in 4 basic Blocks: S, P, D, and F

### Block s elements:-

Diagram illustrating the S-block elements in the periodic table. The S-block elements are highlighted in yellow and include Hydrogen (H), Helium (He), and the elements in groups 1 and 2 (Li, Be, Na, Mg, K, Ca, Rb, Sr, Cs, Ba, Fr, Ra).

- Location: left of the modern periodic table

- It consists of two vertical groups: 1A and 2A

( 1A alkali metals , 2A are called alkaline earth metals )

- All of its elements are solid metals except hydrogen, which is not a gas metal.

### Block p elements:-

Diagram illustrating the P-block elements in the periodic table. The P-block elements are highlighted in yellow and include the elements in groups 13, 14, 15, 16, and 17 (B, C, N, O, F, Ne, Al, Si, P, S, Cl, Ar, Ga, Ge, As, Se, Br, Kr, In, Sn, Sb, Te, I, Xe, Tl, Pb, Bi, Po, At, Rn, Fr, Ra, Ac, Rf, Db, Sg, Bh, Hs, Mt, Ds, Rg, Cr, Nh, Fl, Mc, Lv, Ts, Og).

- Location: Right of the modern periodic table
- It consists of 6 groups starting with group 3A and ending with group zero
- Most of its elements are non-metals, in addition to metalloids and some other metals
- Its elements exist in solid and gaseous forms, with the exception of bromine (Br), which is a liquid nonmetal
- One of its most important groups is the penultimate group 7A, which is called the halogens, and the last group (zero) is called the noble gases.

### Block d Elements

- Location: in the middle of the modern periodic table
- It consists of 10 groups starting with Group 3B and ending with Group 2B
- It begins to appear in the fourth period and its elements are called transitional elements
- All of its elements are metallic elements in the solid state, except for mercury ( Hg ), which is a liquid fluid element

Group	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
1	H												B	C	N	O	F	Ne	
2	Li	Be											Al	Si	P	S	Cl	Ar	
3	Na	Mg											Ga	Ge	As	Se	Br	Kr	
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
6	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn	
7	Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Nh	Fl	Mc	Lv	Ts	Og	
			<b>f-block elements</b>																
			Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu			
			Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr			

### Block f Elements

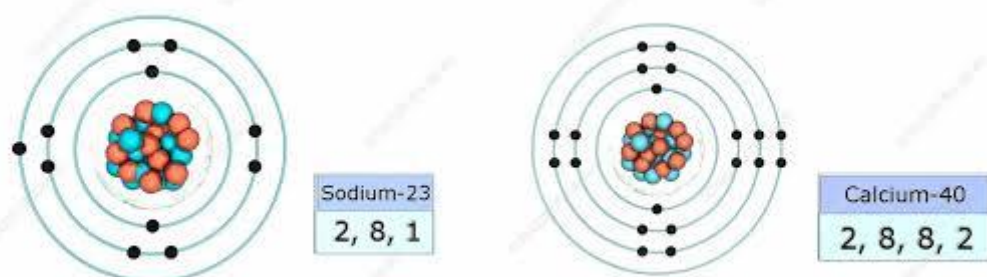
- Location at the bottom of the modern periodic table
- It consists of two series: the lanthanide series and the actinide series.
- All of its elements are metallic elements.

## WEEK 2 : SESSION 1 : UNIT 1 : LESSON 2

**The location of the elements in the groups symbolized by the symbol A can be determined by the atomic number...**

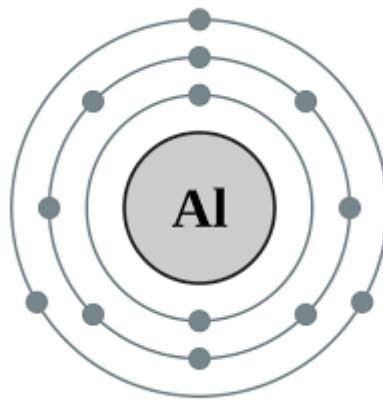
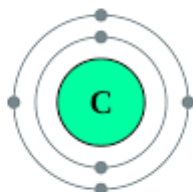
**(1) The electron configuration of the atom is determined ( The number of energy levels is the period number )**

**(2) The number of electrons in the last energy level determines the group number (we put the letter A next to the number of electrons in the last energy level)**



6: Carbon

2,4



$_{11}\text{Na} : 2, 8, 1$

period 3 / group 1A

$_{20}\text{Ca} : 2, 8, 8, 2$

period 4 / group 2A

$_{6}\text{C} : 2, 4$

period 2 / group 4A

$_{13}\text{Al} : 2, 8, 3$

period 3 / group 3A

<b>symbol</b>	<b>Electronic distribution</b>	<b>period</b>	<b>group</b>
${}_{19}\text{K}$			
${}_{8}\text{O}$			
${}_{18}\text{Ar}$			
${}_{17}\text{Cl}$			
${}_{3}\text{Li}$			
${}_{1}\text{H}$			
${}_{4}\text{Be}$			
${}_{9}\text{F}$			

**How to determine the atomic number of an element given its location in the table:-**

**1- The period number of an element is equal to the number of energy levels occupied by the electrons**

**2- The group number of the element is equal to the number of electrons in the last energy level**

**3- The atomic number of an element is equal to the sum of the numbers of electrons in its energy levels**

**Examples:-**

**1- Calculate the atomic number of the element X which is located in the second period and group 1A**

**2- Calculate the atomic number of element Y located in the third period and group 7A**

**3- Calculate the atomic number of the element Z, which is located in the second period and the zero group**

**4- Calculate the atomic number of the element Q, which is located in the third period and group 3A**

**5- Calculate the atomic number of the element M located in the first period and the zero group**

### Elements of one period

- They differ in the number of electrons in the last energy level

- They agree in the number of levels of energy

- They differ in chemical properties

Each element has one more proton than the preceding element in the period

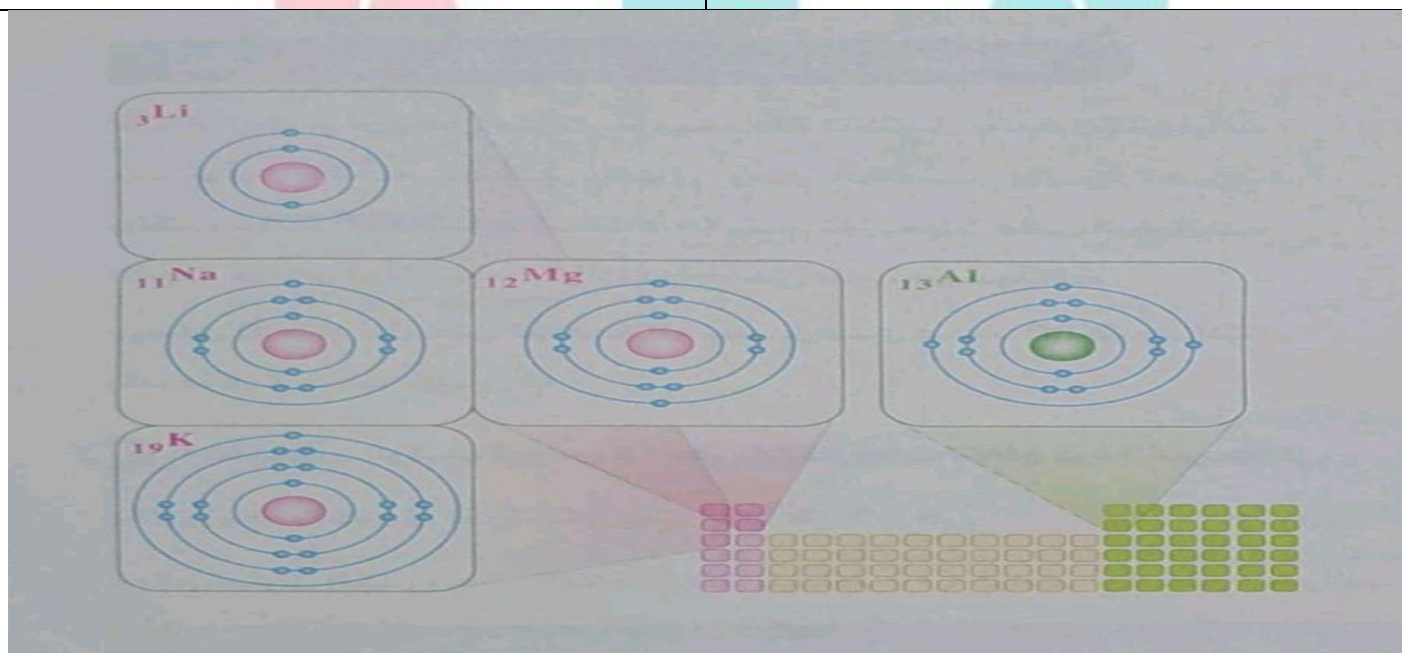
### Elements of one group

- They agree in the number of electrons in the last energy level

- They differ in the number of energy levels

- They are similar in chemical properties

- Each element has 8 electrons more than the previous element, or a full energy level



**Periodic Table of the Elements**

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55 Cs Cesium 132.90545	56 Ba Barium 137.327	57-71 Lanthanide Series	72 Hf Hafnium 178.49	73 Ta Tantalum 180.9479	74 W Tungsten 183.85	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.96655	80 Hg Mercury 200.59	113 Uut Ununtrium unknown	114 Fl Flerovium 289	115 Uup Ununpentium unknown	116 Lv Livermorium 289	117 Uus Ununseptium unknown	118 Uuo Ununoctium unknown
87 Fr Francium 223.0187	88 Ra Radium 226.0254	89-103 Actinide Series	104 Rf Rutherfordium (261)	105 Db Dubnium (262)	106 Sg Seaborgium (263)	107 Bh Bohrium (264)	108 Hs Hassium (265)	109 Mt Meitnerium (266)	110 Ds Darmstadtium (269)	111 Rg Roentgenium (272)	112 Cn Copernicium (285)						
		57 La Lanthanum 138.9055	58 Ce Cerium 140.116	59 Pr Praseodymium 140.90765	60 Nd Neodymium 144.24	61 Pm Promethium 144.9127	62 Sm Samarium 150.36	63 Eu Europium 151.965	64 Gd Gadolinium 157.25	65 Tb Terbium 158.92534	66 Dy Dysprosium 162.50	67 Ho Holmium 164.93032	68 Er Erbium 167.26	69 Tm Thulium 168.93421	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967	
		89 Ac Actinium 227.0278	90 Th Thorium 232.0381	91 Pa Protactinium 231.03626	92 U Uranium 238.02891	93 Np Neptunium 237.0482	94 Pu Plutonium 244.0642	95 Am Americium 243.0614	96 Cm Curium 247.0764	97 Bk Berkelium 247.0763	98 Cf Californium 251.0787	99 Es Einsteinium 252	100 Fm Fermium 257.1037	101 Md Mendelevium 258.1	102 No Nobelium 259.1069	103 Lr Lawrencium 260	
		Alkali Metals	Alkaline Earths	Transition Metals	Basic Metals	Semi-Metals	Nonmetals	Halogens	Noble Gases	Lanthanides	Actinides						

## **WEEK 2 : SESSION 2 : UNIT 1 : LESSON 2**

**The elements are divided according to their properties and the distribution of electrons into four types of elements:-**

**Metals                      nonmetals                      metalloids                      inert gases.**

**Metals:-**

**They are characterized by the last energy level often contains less than four electrons**

**Nonmetals:-**

**They are characterized by the last energy level often contains more than four electrons**

**Metalloids:-**

**elements that combine the properties of metals and nonmetals. These elements fall into the S block.**

**Inert gases:-**

**gaseous elements that do not react under normal conditions because the last energy level is filled with electrons**

**Types of gaseous elements:-**

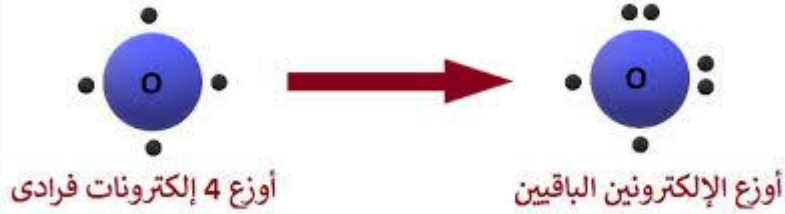
<b><u>Inert gases</u></b>	<b><u>Gases (non-metallic)</u></b>
<b><u>- It is in the zero group</u></b>	<b><u>It is found in the P-block elements, except for hydrogen (H), which is in the S-block</u></b>
<b><u>- 6 items</u></b>	<b><u>- 5 items</u></b>
<b><u>- Do not participate in chemical reactions under normal conditions</u></b>	<b><u>- Participates in chemical reactions under normal conditions</u></b>
<b><u>- It includes the elements Helium(He), Neon(Ne), Argon(Ar), Krypton(Kr), Xenon(Xe), and Radion(Rn)</u></b>	<b><u>-It includes the elements Hydrogen(H) – Nitrogen(N)- Oxygen(O) – Fluorine(F) – Chlorine(Cl)</u></b>



## The electrons of the last energy level play an important role in the formation of chemical bonds

### valence:-


It is the number of single electrons in the Lewis structure of an element



### مثال توضيحي

التمثيل النقطي لإلكترونات تكافؤ ذرة الأكسجين O

- التوزيع الإلكتروني لذرة الأكسجين هو
- طبقاً لطريقة لويس يتم توزيع إلكترونات المستوى الخارجي فرادي أولاً على الجوانب الأربعة لرمز العنصر، ثم يبدأ الأزواج حتى يتم توزيعها جميعاً كالتالي:



**التكافؤ (Valency):**  
هو عدد الإلكترونات المفردة في تركيب لويس للعنصر.

الجدول الآتي يوضح التمثيل النقطي بطريقة لويس لإلكترونات مستوى الطاقة الأخير (إلكترونات التكافؤ) لمجموعات الجدول الدوري من المجموعة 1A حتى المجموعة الصفيرية.

رقم المجموعة	1A	2A	3A	4A	5A	6A	7A	0
العنصر	Li	Be	B	C	N	O	F	Ne
التكافؤ	أحادي	ثنائي	ثلاثي	رباعي	ثلاثي	ثنائي	أحادي	0

الجدول التالي يوضح بعض المعلومات الخاصة بعناصر الدورة (2) من الجدول الدوري الحديث بدلالة تركيب لويس النقطي:

عناصر الدورة (2)	${}^3\text{Li}$	${}^4\text{Be}$	${}^5\text{B}$	${}^6\text{C}$	${}^7\text{N}$	${}^8\text{O}$	${}^9\text{F}$	${}^{10}\text{Ne}$
التوزيع الإلكتروني	2, 1	2, 2	2, 3	2, 4	2, 5	2, 6	2, 7	2, 8
تركيب لويس النقطي	Li	Be	B	C	N	O	F	Ne
التكافؤ بمعلومية تركيب لويس	أحادي	ثنائي	ثلاثي	رباعي	ثلاثي	ثنائي	أحادي	صفر
رقم مجموعة العنصر	1A	2A	3A	4A	5A	6A	7A	0

ويلاحظ من الجدول السابق أن:

تكافؤ عناصر المجموعات من 1A : 4A يساوي رقم المجموعة.

تكافؤ عناصر المجموعات من 5A : 0 يساوي (8 - عدد إلكترونات مستوى الطاقة الأخير).

تكافؤ مجموعة الغازات الخاملة يساوي صفر... **علل؟**  
لاكتمال مستوى الطاقة الأخير في ذراتها بالإلكترونات وبالتالي لا يحتوي تركيب لويس لها على إلكترونات مفردة.

Atomic number	Name of element	Symbol	Valency	Atomic number	Name of element	Symbol	Valency
1	Hydrogen	H	1	11	Sodium	Na	1
2	Helium	He	0	12	Magnesium	Mg	2
3	Lithium	Li	1	13	Aluminium	Al	3
4	Beryllium	Be	2	14	Silicon	Si	4
5	Boron	B	3	15	Phosphorous	P	3
6	Carbon	C	4	16	Sulphur	S	2
7	Nitrogen	N	3	17	Chlorine	Cl	1
8	Oxygen	O	2	18	Argon	Ar	0
9	Fluorine	F	1	19	Potassium	K	1
10	Neon	Ne	0	20	Calcium	Ca	2

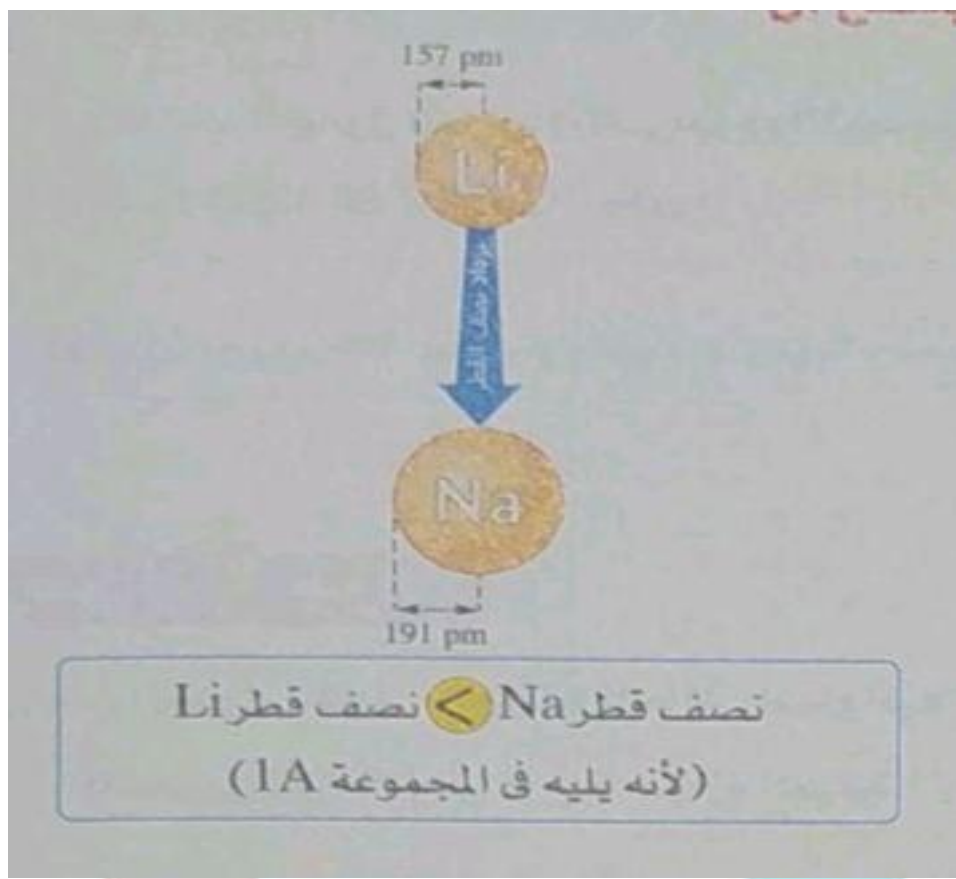
**The electron configuration of element atoms reflects their properties**

**The chemical properties of elements depend on the number of electrons in the last energy level**

**- Graduation of the physical properties of some alkali metals and halogens**

والجدولان التاليان يوضحان تدرج بعض الخواص الفيزيائية لبعض فلزات الأقلء والهالوجينات :

درجة الغليان	درجة الانصهار	نصف قطر الذرة	الهالوجينات	درجة الغليان	درجة الانصهار	نصف قطر الذرة	فلزات الأقلء
-34°C	-101°C	99 pm	كلور $^{17}\text{Cl}$ 2, 8, 7	1347°C	181°C	157 pm	ليثيوم $^3\text{Li}$ 2, 1
59°C	-7°C	114 pm	بروم $^{35}\text{Br}$ 2, 8, 18, 7	883°C	98°C	191 pm	صوديوم $^{11}\text{Na}$ 2, 8, 1
184°C	114°C	133 pm	يود $^{53}\text{I}$ 2, 8, 18, 18, 7	774°C	64°C	235 pm	بوتاسيوم $^{19}\text{K}$ 2, 8, 8, 1
تزداد ↑	تزداد ↑	يزداد ↑		تقل ↓	تقل ↓	يزداد ↑	
بزيادة العدد الذري				بزيادة العدد الذري			



**The relationship between the physical state of an element and its melting and boiling points at room temperature:-**

<u>The physical state of matter</u>	<u>Melting point</u>	<u>Boiling point</u>
<u>Solid</u> elements	<u>more than 25</u>	
<u>Liquid</u> elements	<u>less than 25</u>	<u>greater than 25</u>
<u>Gaseous</u> elements	<u>less than 25</u>	

**Give a reason:** The melting and boiling points of lithium(Li) and potassium(K) are higher than room temperature?

Because both of them are solid elements at room temperature.

**Give a reason:** The melting and boiling points of chlorine(Cl) are lower than room temperature?

Because chlorine is a gaseous element At room temperature

## Chemical activity

1) In the two groups of alkali(1A) and alkali earth metals(2A), the degree of activity increases with an increase in the atomic number as it descends downward in the group.

2) The degree of activity of the alkali metal elements is greater than the degree of activity of the ground alkali metals

3) In the group of Halogens(7A), the chemical activity decreases as the atomic number increases from top to bottom

Sizes of atoms and their ions in pm

Group 1		Group 2		Group 3		Group 4		Group 5	
Li <sup>+</sup> 90	Li 134	Li <sup>+</sup> 90	Be 134	B <sup>3+</sup> 41	B 82	O 73	O <sup>2-</sup> 126	F 71	F <sup>-</sup> 119
Na <sup>+</sup> 116	Na 154	Mg <sup>2+</sup> 90	Mg 130	Al <sup>3+</sup> 41	Al 82	S 102	S <sup>2-</sup> 170	Cl 99	Cl <sup>-</sup> 167
K <sup>+</sup> 152	K 196	Ca <sup>2+</sup> 114	Ca 174	Ga <sup>3+</sup> 76	Ga 126	Se 116	Se <sup>2-</sup> 184	Br 114	Br <sup>-</sup> 182
Rb <sup>+</sup> 166	Rb 211	Sr <sup>2+</sup> 132	Sr 192	In <sup>3+</sup> 94	In 144	Te 135	Te <sup>2-</sup> 207	I 133	I <sup>-</sup> 206

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## Exercises : Complete the following phrases:-

1. The period number is equal to the number of -----
2. Block s is located ----- the modern periodic table
3. Block p contains ----- group
4. Noble gases fall into the category -----
5. Block d is located in ----- the modern periodic table
6. The first scientist to create a true periodic table -----
7. Moseley arranged the elements in ascending order according to -----
8. Number of elements in the modern periodic table -----
9. The scientist discovered ----- that the nucleus contains positively charged protons
10. The scientist added ----- the group of inert gases to the periodic table
11. The Block ----- contains the largest number of items
12. The alkali earth group is located in -----the periodic table

13. The zero group in the modern periodic table belongs to the block  
-----
14. Potassium belongs to the group -----
15. The 1A group is called -----, the 2A group is called -----,  
the 7A group is called -----, and the zero group is called-- -----
16. The transitional elements begin to appear in the period -----
17. The valency of the alkali group elements ----- while the  
halogen group -----
18. The element that falls in the third period and group 3A has the  
atomic number -----
19. In group 1A, it increases ----- from top to bottom
20. In halogens, the boiling point ----- from top to bottom, in  
alkali elements, the melting point ----- from top to bottom.